# **SULPHIDE ION**



### 1. PERFORMANCE

1) Sampling method : Immersion method (Refer to Page 17)
2) Measuring range : 0.5-10 ppm

2) Measuring range  $1.5 \times 10^{\circ}$  ppm  $1.5 \times 10^{\circ}$  Sample volume  $1.5 \times 10^{\circ}$  Cover  $1.5 \times 10^{\circ}$  Measurement  $1.5 \times 10^{\circ}$  Carbon shall be a substituting time  $1.5 \times 10^{\circ}$  Carbon shall be a substituti

9) Reading : Direct reading from the scale 10) Colour change : White → Pale brown

### 2. RELATIVE STANDARD DEVIATION

RSD-low: 10% RSD-mid.: 5% RSD-high: 5%

## 3. CHEMICAL REACTION

By reacting with Lead acetate, Lead sulphide is produced.  $S^2 - + (CH_3CO_2)_2Pb \rightarrow PbS$ 

## 4. CALIBRATION OF THE TUBE

SODIUM SULPHIDE STANDARD SOLUTION METHOD

### 5. INTERFERENCE AND CROSS SENSITIVITY

Substance	Interference	ppm	Coexistence
Carbonate ion		1,000	Lower readings are given.
Chloride ion		2,000	Higher readings are given.
Copper ion			The accuracy of readings is not affected.

## 6. SAMPLING METHOD

(Immersion method)

- 1) Cut both ends of a fresh detector tube with a file.
- 2) Immerse the filled end of the tube with white end plug into the shall prepared sample solution. Capillary action will occur immediately and the sample solution shall rise through the reagent. If Chloride ion is existed in the sample solution, a discolouration will be occurred in the detecting reagent layer from its inlet and the discoloured layer shall be given according to the concentration of Chloride ion.

